

# Naming Compounds Review

Fill in the blank with the correct word. Some words may be used more than once.

cations

electroneutrality

*-ide*

Roman numerals

polyatomic

ionic

subscript

1. Equal amounts of positive and negative charges is called \_\_\_\_\_.
2. A charged group of two or more bonded atoms that acts like a single ion is a \_\_\_\_\_ ion.
3. \_\_\_\_\_ are never found without anions (or sometimes electrons) nearby to neutralize the charges.
4. The names for \_\_\_\_\_ come from the element from which they are formed.
5. When the same element forms two or more positive ions, \_\_\_\_\_ are used to tell the difference between the ions.
6. The name of a negative ion is made by changing the ending of the element name to \_\_\_\_\_.
7. A \_\_\_\_\_ is a number written below and to the right of an element's symbol; it tells the number of atoms in a formula.
8. If a chemical compound is composed of oppositely charged ions is called a(n) \_\_\_\_\_ compound.

**Write the formula for the following compounds.**

9. potassium bromide

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10. barium fluoride

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11. tin(IV) oxide

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12. cesium bromide

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13. cobalt(II) bromide

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14. mercury(I) sulfide

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15. aluminum iodide

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**Write the names of the following, and include Roman numerals if needed.**

16.  $\text{Cu}^+$  \_\_\_\_\_

24.  $\text{Fe}_2\text{O}_3$  \_\_\_\_\_

17.  $\text{Cl}^-$  \_\_\_\_\_

25.  $\text{Mg}_3\text{N}_2$  \_\_\_\_\_

18.  $\text{O}^{2-}$  \_\_\_\_\_

26.  $\text{ZnO}$  \_\_\_\_\_

19.  $\text{P}^{3-}$  \_\_\_\_\_

27.  $\text{NaI}$  \_\_\_\_\_

20.  $\text{Na}^+$  \_\_\_\_\_

28.  $\text{Cu}_2\text{O}$  \_\_\_\_\_

21.  $\text{Mg}^{2+}$  \_\_\_\_\_

29.  $\text{CaCl}_2$  \_\_\_\_\_

22.  $\text{NaCl}$  \_\_\_\_\_

30.  $\text{CrCl}_3$  \_\_\_\_\_

23.  $\text{ZnS}$  \_\_\_\_\_

31.  $\text{HgO}$  \_\_\_\_\_

**Write the name for the following compounds that contain polyatomic ions.**

**38.**  $K_2Cr_2O_7$  \_\_\_\_\_

**39.**  $KClO_2$  \_\_\_\_\_

**40.**  $Fe(ClO_3)_3$  \_\_\_\_\_

**41.**  $Na_2SO_4$  \_\_\_\_\_

**42.**  $Na_2SO_3$  \_\_\_\_\_

**43.**  $KMnO_4$  \_\_\_\_\_

**44.**  $Pb(NO_3)_2$  \_\_\_\_\_

**Write the formula for the following compounds containing polyatomic ions.**

**45.** lead (IV) chromate \_\_\_\_\_

**46.** sodium chlorite \_\_\_\_\_

**47.** magnesium nitrate \_\_\_\_\_

**48.** sodium nitrite \_\_\_\_\_

**49.** hydrogen cyanide \_\_\_\_\_

**50.** aluminum hydroxide \_\_\_\_\_

**51.** ammonium sulfate \_\_\_\_\_