

Physical Science

Chapter 16 Reading Quiz

Name: _____

1. What is the second assumption of the kinetic theory?

2. When the temperature of a substance is lowered, the particles will have less thermal energy and will vibrate more _____.
3. In science _____ means the average kinetic energy of particles in a substance.
4. _____ do not have a definite volume or shape.
5. The particles of a _____ are tightly packed together and are constantly vibrating in place.
6. What is the temperature at which a solid begins to liquefy? _____
7. What is it called when the vapor pressure of a liquid is equal to the external pressure acting on the surface of the liquid? _____
8. What is the normal boiling point of water in degrees Celsius? _____
9. In what state of matter are the electrons stripped from the atom?

10. What types of materials do not have a definite melting point? _____
11. _____ forms when lava, made of molten rock, cools quickly as when it spills into water.
12. Liquid crystals are used in the display of calculators because they respond to electric _____.
13. _____ is the ability of a fluid to exert an upward force on an object immersed in it.
14. _____ found that the buoyant force on an object is equal to the weight of the fluid displaced by the object.
15. What does most of a ship's hull contain? _____
16. A liquid's resistance to flow is called _____.
17. Who came up with the idea that as the velocity of a fluid increases, the pressure decreases? _____
18. How many Newtons per square meter are exerted at the surface of the earth?

19. Which part of the atmosphere do the auroras form in? _____
20. What happens to a balloon as it rises through the atmosphere?

21. According to figure 21, what is the pressure at 300 L? _____
22. What states that as pressure decreases, volume increases _____
23. Write the equation that is used for pressure and volume relationships of a gas.

24. What states that as the temperature of a gas is lowered, the volume decreases?

25. Write the equation that is used for temperature and volume relationships of a gas.
