1. Determine the number of atoms present in 4.00 mol of aluminum.

2. Determine the number of atoms present in 1.55 mol of sodium.

3. Convert 2.65×10^{25} atoms of fluorine to moles of fluorine atoms.

4. Convert 4.26×10^{25} molecules of hydrogen, H_2, to moles of hydrogen, H_2.

5. Convert 1.75×10^{26} atoms of potassium to moles of potassium.

6. Determine the mass in grams of 7.20 mol of antimony.

7. Determine the mass in grams of 0.500 mol of uranium.

8. Determine the mass in grams of 0.750 mol of francium.

9. A sample of lead has a mass of 150.0 g. What amount of lead in moles does the sample contain?

10. A sample of gold has a mass of 5.00×10^{-3} g. What amount of gold in moles does the sample contain?